

CLASS 11

CHEMISTRY

REVISION NOTES

Class 11 Chemistry Chapter 4 Chemical Bonding and Molecular Structure – Revision Notes

Quick revision notes for Class 11 Chemistry Chapter 4 (Chemical Bonding and Molecular Structure). These Class 11 Chemistry Chapter 4 notes are ideal for quick revision just before exams.

Contents

- Bond Types
- Shape and Hybridization
- Molecular Orbital Theory
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Bond Types

Ionic: electron transfer.

Covalent: electron sharing; octet rule.

Shape and Hybridization

VSEPR theory predicts shape via electron pair repulsion.

sp (linear), sp² (trigonal planar), sp³ (tetrahedral).

Molecular Orbital Theory

Bond order = $\frac{1}{2}(\text{bonding} - \text{antibonding electrons})$.

Hydrogen Bonding

Between H (bonded to N/O/F) and an electronegative atom; explains water's high boiling point.

One-Line Summary

Chemical bonding theories (ionic, covalent, VSEPR, hybridization, and molecular orbital theory) explain how and why atoms combine, and predict the shape, stability, and properties of the resulting molecules.

Class 11 Chemistry Chapter 4 – Solutions and Important Questions

Need full answers or more practice? See the [Class 11 Chemistry Chapter 4 Solutions](#) and [Class 11 Chemistry Chapter 4 Extra Questions](#).

See also: [Chapter 1](#) | [Chapter 2](#) | [Chapter 3](#) | [Chapter 4](#)

Practice more: [Chapter 1](#) | [Chapter 2](#) | [Chapter 3](#) | [Chapter 4](#)

Quick revision: [Chapter 1](#) | [Chapter 2](#) | [Chapter 3](#)