

CLASS 11

CHEMISTRY

REVISION NOTES

## Class 11 Chemistry Chapter 5 Chemical Thermodynamics – Revision Notes

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Quick revision notes for Class 11 Chemistry Chapter 5 (Chemical Thermodynamics). These Class 11 Chemistry Chapter 5 notes are ideal for quick revision just before exams.

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Quick revision notes for Class 11 Chemistry Chapter 5 (Chemical Thermodynamics). These Class 11 Chemistry Chapter 5 notes are ideal for quick revision just before exams.

## Systems and First Law

Open, closed, isolated systems.

$$\Delta U = q + w.$$

## Enthalpy

$$H = U + PV; \Delta H = q_p.$$

Exothermic:  $\Delta H < 0$ ; Endothermic:  $\Delta H > 0$ .

Hess's law:  $\Delta H$  is path-independent (state function).

## Entropy and Free Energy

Entropy: measure of disorder; increases in spontaneous processes (isolated system).

$$\Delta G = \Delta H - T\Delta S; \Delta G < 0 \text{ spontaneous, } \Delta G = 0 \text{ equilibrium.}$$

## One-Line Summary

Chemical thermodynamics uses enthalpy, entropy, and Gibbs free energy to determine whether a chemical process will occur spontaneously under given conditions.

## Class 11 Chemistry Chapter 5 – Solutions and Important Questions

Need full answers or more practice? See the [Class 11 Chemistry Chapter 5 Solutions](#) and [Class 11 Chemistry Chapter 5 Extra Questions](#).

**See also:** [Chapter 1](#) | [Chapter 2](#) | [Chapter 3](#) | [Chapter 4](#) | [Chapter 5](#)

**Practice more:** [Chapter 1](#) | [Chapter 2](#) | [Chapter 3](#) | [Chapter 4](#) | [Chapter 5](#)

**Quick revision:** [Chapter 1](#) | [Chapter 2](#) | [Chapter 3](#) | [Chapter 4](#)