

CLASS 11

CHEMISTRY

REVISION NOTES

Class 11 Chemistry Chapter 6 Equilibrium – Revision Notes

Quick revision notes for Class 11 Chemistry Chapter 6 (Equilibrium). These Class 11 Chemistry Chapter 6 notes are ideal for quick revision just before exams.

Contents

Equilibrium Basics

Le Chatelier's Principle

Acids, Bases, pH

Buffers

One-Line Summary

Class 11 Chemistry Chapter 6 – Solutions and Important Questions

Quick revision notes for Class 11 Chemistry Chapter 6 (Equilibrium). These Class 11 Chemistry Chapter 6 notes are ideal for quick revision just before exams.

Equilibrium Basics

Dynamic equilibrium: forward rate=reverse rate.

$K_c = \frac{[\text{products}]}{[\text{reactants}]}$ (each raised to stoichiometric coefficient).

Le Chatelier's Principle

Equilibrium shifts to counteract imposed change (concentration, pressure, temperature).

Acids, Bases, pH

Arrhenius, Brønsted-Lowry, Lewis definitions.

$\text{pH} = -\log[\text{H}^+]$; $K_w = 10^{-14}$ at 25°C .

Buffers

Resist pH change; weak acid/base + conjugate; common ion effect.

One-Line Summary

Equilibrium describes the dynamic balance between forward and reverse reactions, governed by the equilibrium constant and Le Chatelier's principle, extending to acid-base and buffer systems in solution.

Class 11 Chemistry Chapter 6 – Solutions and Important Questions

Need full answers or more practice? See the [Class 11 Chemistry Chapter 6 Solutions](#) and [Class 11 Chemistry Chapter 6 Extra Questions](#).

See also: [Chapter 1](#) | [Chapter 2](#) | [Chapter 3](#) | [Chapter 4](#) | [Chapter 5](#) | [Chapter 6](#)

Practice more: [Chapter 1](#) | [Chapter 2](#) | [Chapter 3](#) | [Chapter 4](#) | [Chapter 5](#) | [Chapter 6](#)

Quick revision: [Chapter 1](#) | [Chapter 2](#) | [Chapter 3](#) | [Chapter 4](#) | [Chapter 5](#)