

CLASS 11

CHEMISTRY

REVISION NOTES

Class 11 Chemistry Chapter 7 Redox Reactions – Revision Notes

Quick revision notes for Class 11 Chemistry Chapter 7 (Redox Reactions). These Class 11 Chemistry Chapter 7 notes are ideal for quick revision just before exams.

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Quick revision notes for Class 11 Chemistry Chapter 7 (Redox Reactions). These Class 11 Chemistry Chapter 7 notes are ideal for quick revision just before exams.

Oxidation and Reduction

Oxidation: loss of electrons (oxidation number increases).

Reduction: gain of electrons (oxidation number decreases).

Oxidation Number Rules

Free element=0; O usually -2 ; H usually $+1$.

Agents and Types

Oxidizing agent: gets reduced. Reducing agent: gets oxidized.

Disproportionation: same element oxidized and reduced.

Balancing Methods

Oxidation number method; ion-electron (half-reaction) method.

One-Line Summary

Redox reactions involve simultaneous oxidation (electron loss) and reduction (electron gain), tracked using oxidation numbers and balanced via the oxidation number or half-reaction methods.

Class 11 Chemistry Chapter 7 – Solutions and Important Questions

Need full answers or more practice? See the [Class 11 Chemistry Chapter 7 Solutions](#) and [Class 11 Chemistry Chapter 7 Extra Questions](#).

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