

CLASS 7

SCIENCE

REVISION NOTES

Revision Notes for Class 7 Science Chapter 4: The World of Metals and Non-metals

Quick revision notes summarizing all key concepts from Class 7 Science Chapter 4 "The World of Metals and Non-metals".

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Metals are generally lustrous, malleable (can be beaten into sheets), ductile (can be drawn into wires), sonorous (produce a ringing sound), and good conductors of heat and electricity. Examples: copper, aluminium, iron, gold, zinc, sodium.

Non-metals are generally non-lustrous, brittle, poor conductors of heat and electricity, and often exist as gases at room temperature. Examples: sulphur, oxygen, nitrogen, carbon, phosphorus. Exception: graphite (a form of carbon) is a non-metal but conducts electricity well.

Metal oxides are generally basic in nature (turn red litmus paper blue), e.g., magnesium oxide.

Non-metal oxides are generally acidic in nature (turn blue litmus paper red), e.g., sulphur dioxide forms sulphurous acid in water.

Rusting of iron requires both air (oxygen) and water/moisture to occur; removing either factor prevents rusting.

Sodium is a highly reactive metal that reacts vigorously with water/air and is stored in kerosene oil for safety.

Copper is preferred over iron for boiling-water vessels because it conducts heat well and resists rusting/corrosion better.

Only a few metals (like gold and silver) are used for jewellery because they are lustrous, soft enough to shape, and resist corrosion.

Protective coatings (like zinc galvanisation) prevent iron from rusting by blocking air and moisture contact.

Everyday uses of metals: cooking vessels (heat conduction), electrical wiring (electrical conduction, ductility), foil/wrapping (malleability). Everyday uses of non-metals: insulation (poor conductors), matchsticks (brittle, reactive), breathing (oxygen), pencils/batteries (graphite).

NCERT Solutions: [Class 7 Science Chapter 4](#)

Extra Questions: [Class 7 Science Chapter 4](#)

For the full chapter, see the [Class 7 Science book page](#) and the [Class 7 Science Formulas Handbook](#).